

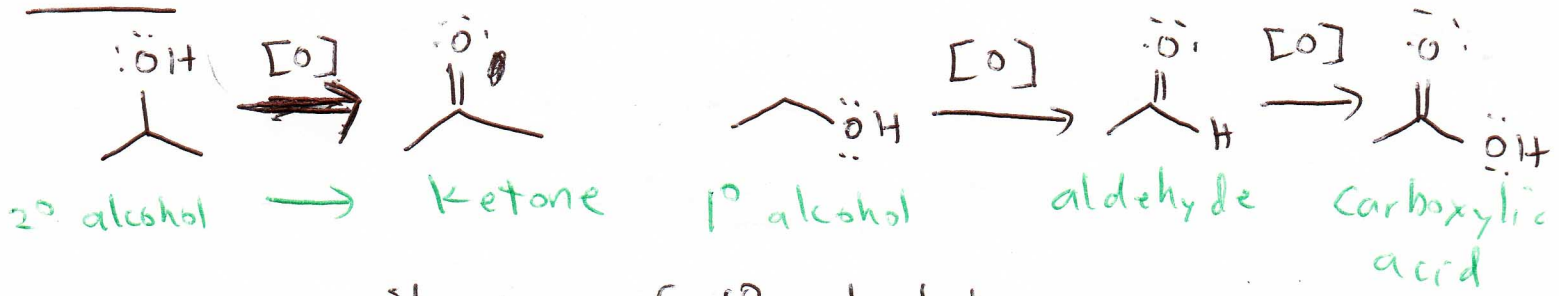
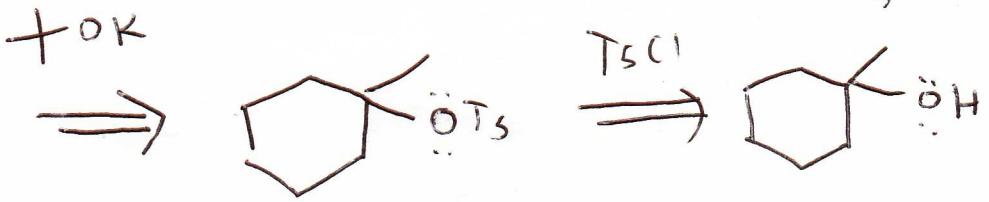
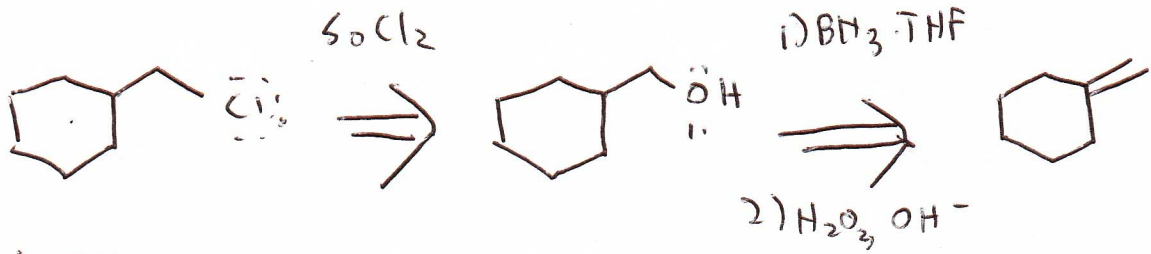
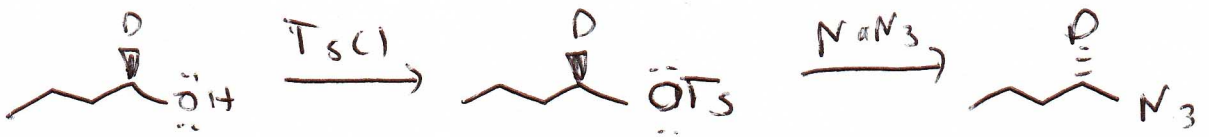
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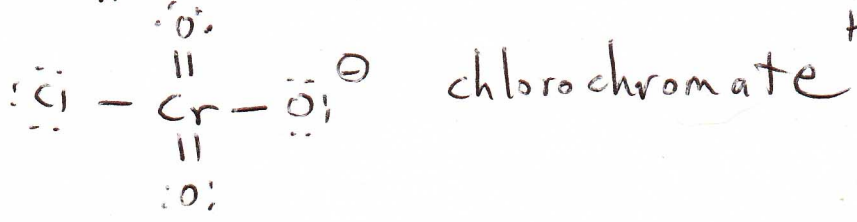
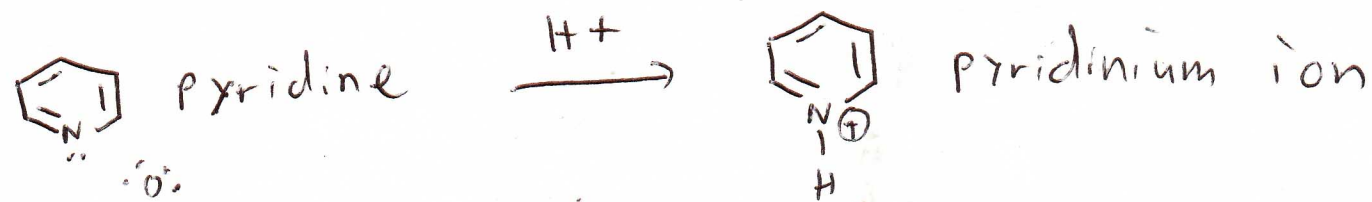
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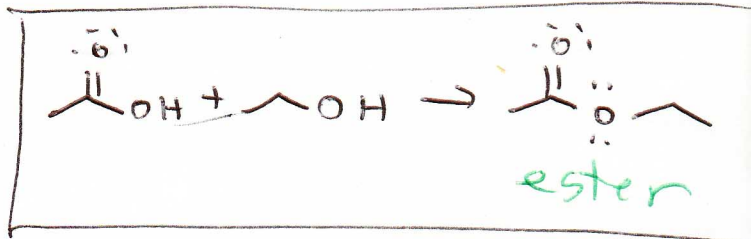
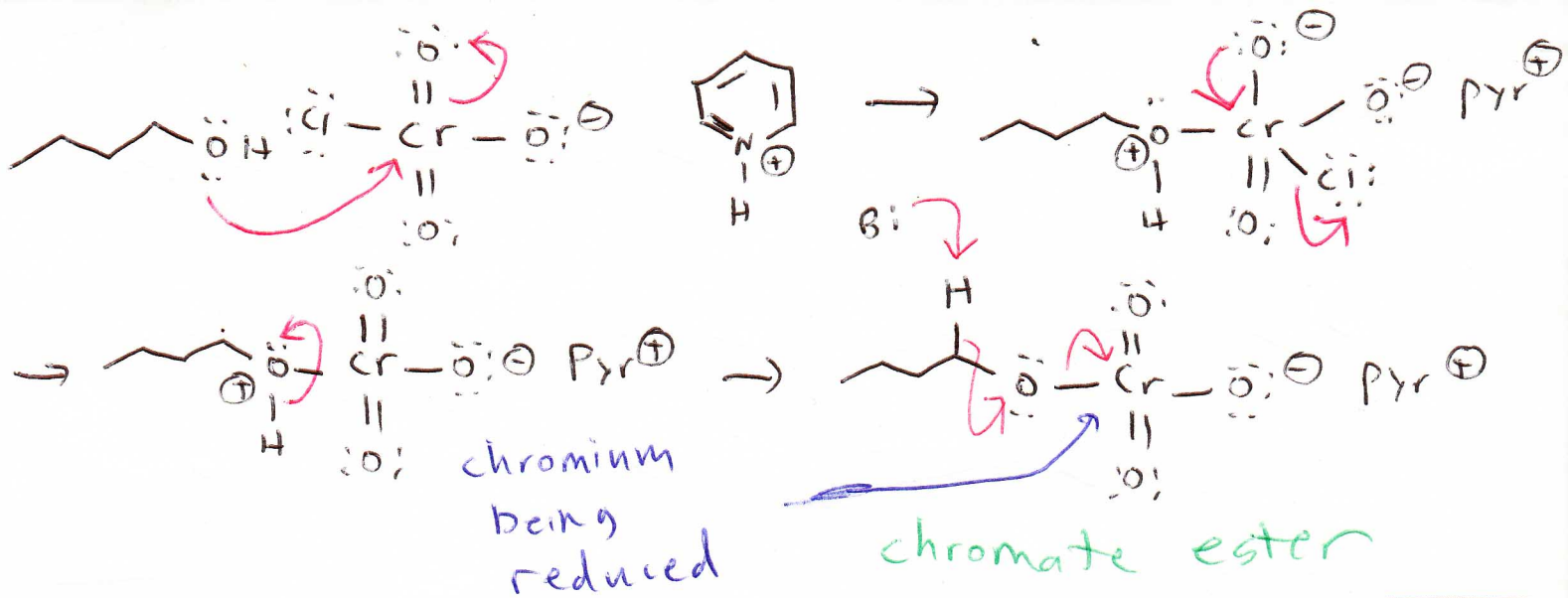
wrong configuration



Selective oxidation of 1° alcohols

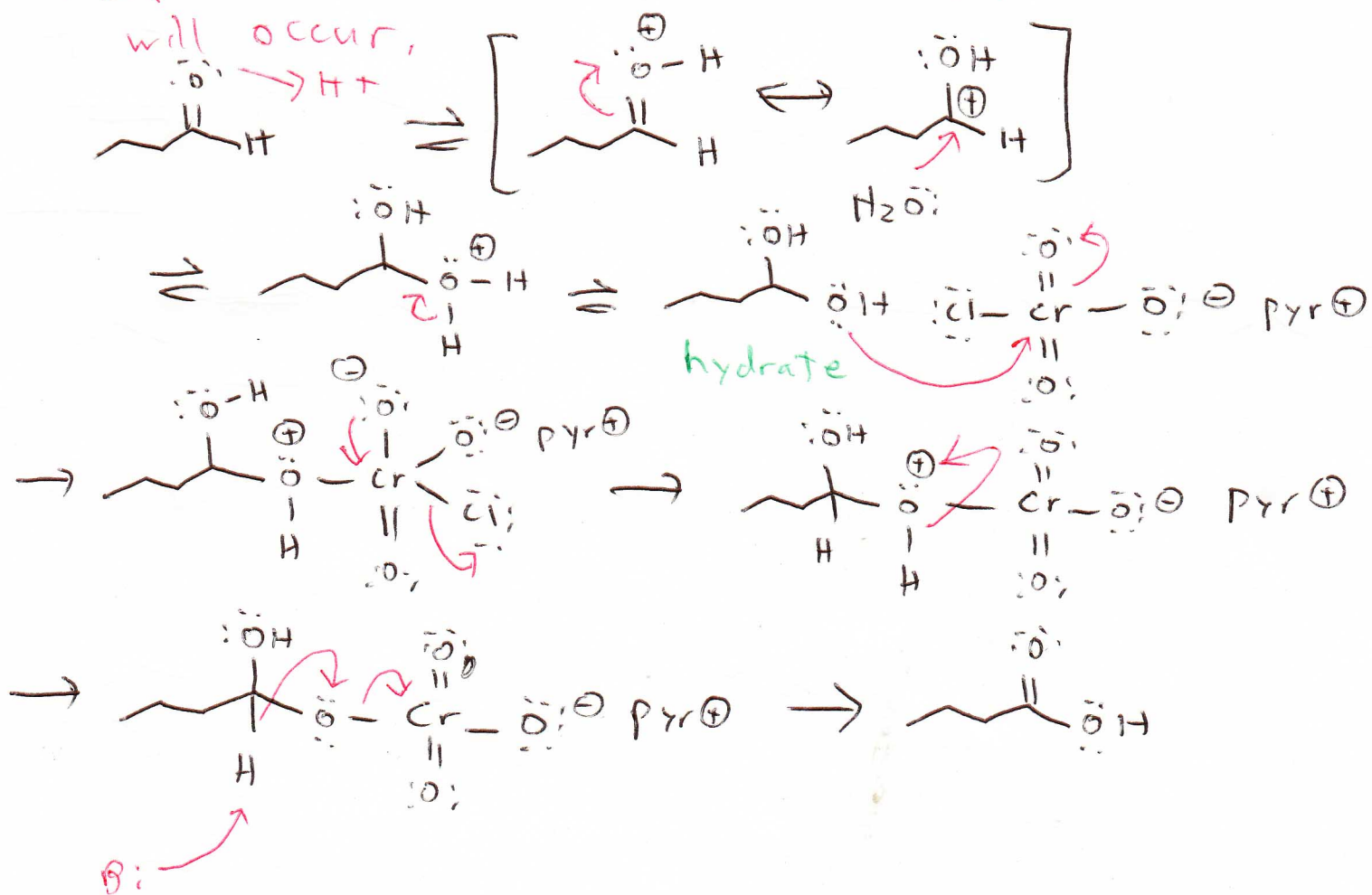
PCC ~~pyridinium~~ pyridinium chlorochromate



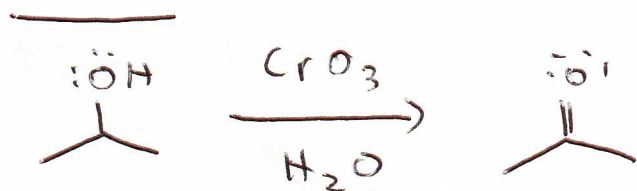
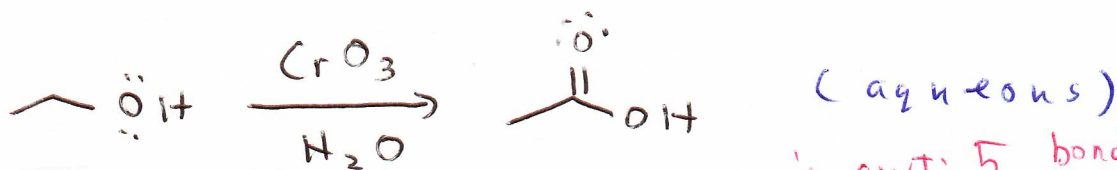
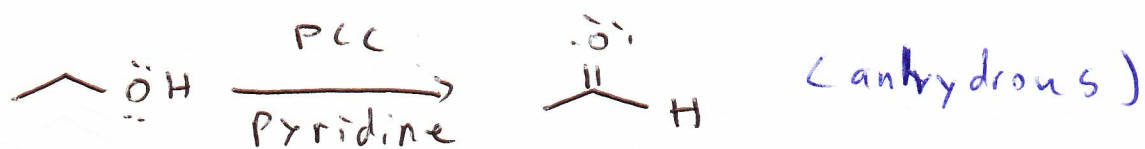


PCC can be used in anhydrous conditions to make an aldehyde.

\*If there is any water around, "over" oxidation will occur.



Formation of a hydrate is often thermodynamically unfavorable. However, the formation is a reversible equilibrium process, and if it can happen (if water is present), it will happen, even if only a trace of hydrate forms at any one time. Once the hydrate forms (even briefly), it can become oxidized, which is why oxidation of 1° alcohols can lead to the formation of carboxylic acids.



can't exist; 5 bonds to carbon

(no carboxylic acid)

