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electron pair geometry — the geometry that a molecule will adopt ignoring the difference between bonding and lone pairs.

<u>total # of electron pairs</u>	<u># lone pairs</u>		<u>electron pair geometry</u>	<u>molecular geometry</u>
4	0		tetrahedral	tetrahedral
4	1		tetrahedral	trigonal pyramidal
4	2		tetrahedral	bent
3	0		trigonal planar	trigonal planar
3	1		trigonal planar	bent
2	0		linear	linear

$H \leftarrow Be \rightarrow H$ molecules will be polar if:

- 1) there are lone pairs around the central atom (true for all shapes with 4 or less electron pairs)
- 2) the atoms attached to the center atom are not all the same

