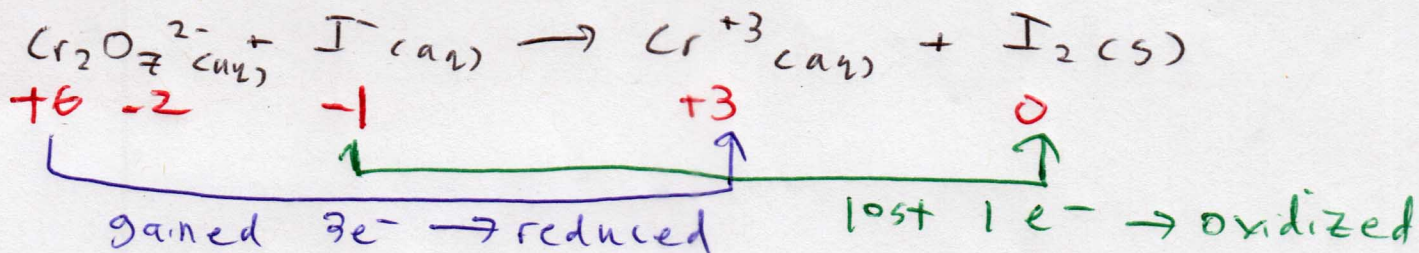


2/18/15

Quiz #3 - Monday 2/23

Lab Exam #1 - Wednesday 2/25

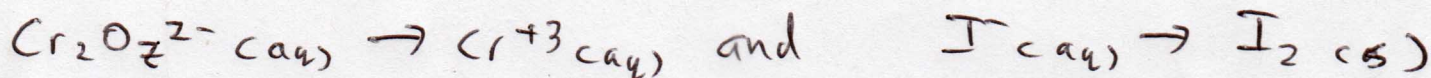
Exam #3 - Monday 3/2



The oxidizing agent gets reduced $\rightarrow \text{Cr}_2\text{O}_7^{2-}$

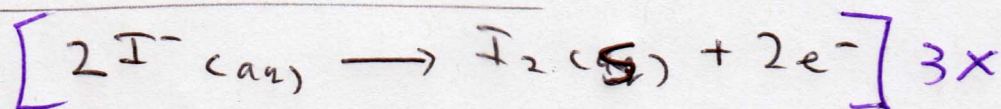
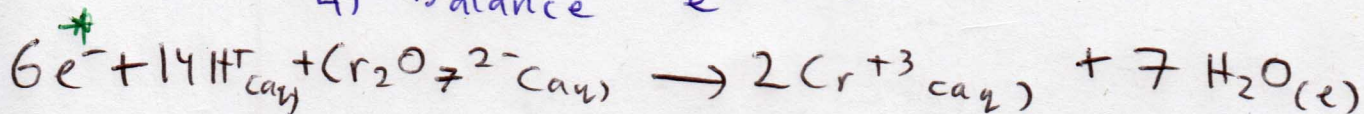
The reducing agent gets oxidized $\rightarrow \text{I}^-$

Half-reaction: Either the oxidation or reduction portion of a redox reaction

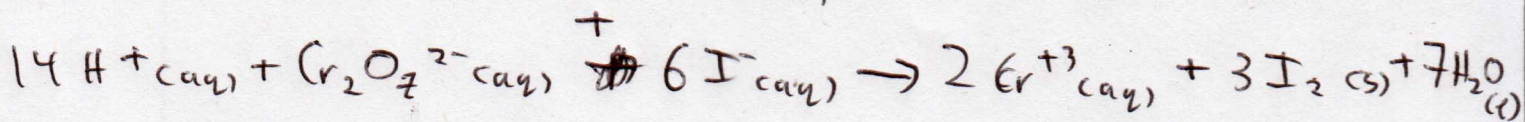
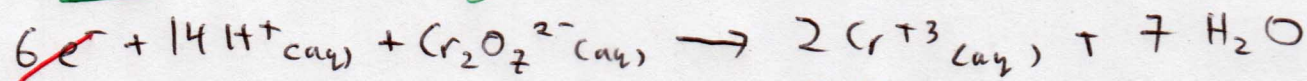


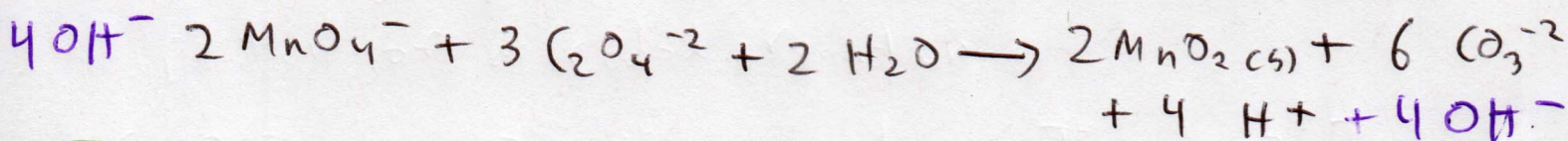
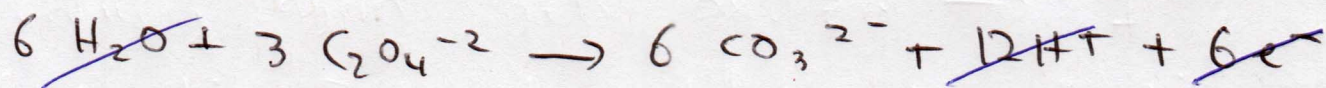
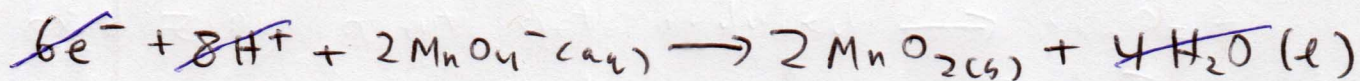
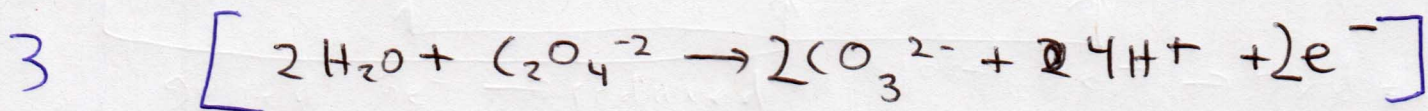
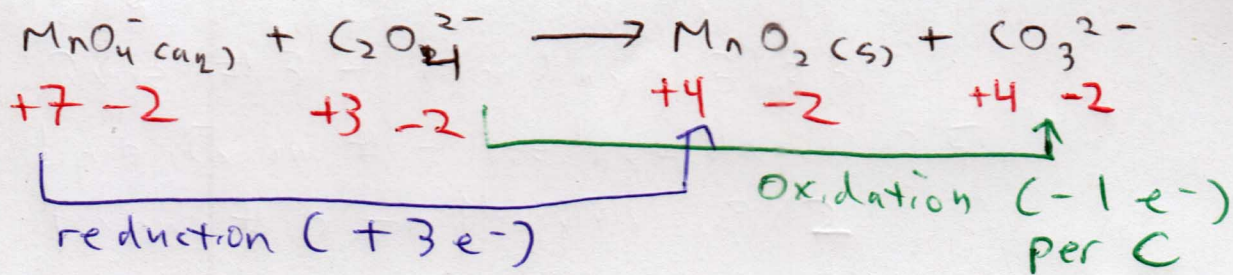
Balancing in acidic conditions

- 1) balance atoms other than O or H
- 2) balance O using H_2O
- 3) balance H using H^+
- 4) balance e^-



* These half-reactions are balanced individually, but the total # of e^- gained or lost must match,





For basic conditions, add one OH⁻ to each side for each H⁺ present

